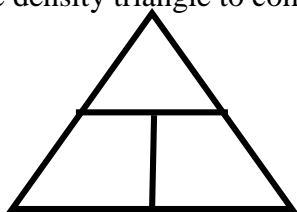


Name: _____ Due Date: _____ Period: _____

Chapter 2, section 2 assessment

1. Define Mass
2. Why is mass more useful than weight for measuring matter?
3. Define volume
4. Define density
5. What measurements must you make to determine the density of a sample of matter?
6. How can you determine whether a solid substance is more or less dense than water?

Use the density triangle to complete the following density calculations.



7. A piece of metal has a volume of 38cm^3 and a mass of 277g . Calculate its density.
8. Iron has a density of 7.9 g/cm^3 . Calculate the mass of 38 cm^3 of Iron.
9. Lead has a density of 11.3 g/cm^3 . What would be the volume of 277 g . of Lead?
10. Water has a density of 1g/ml . Rubbing alcohol has a density of 0.8 g/ml . Honey has a density of 1.4 g/ml . milk has a density of 1.2 g/ml . If you were to carefully pour all four of these into a glass, they would settle into layers. Sketch and label the layers.

