

Name: _____ Due Date: _____ Period: _____

Chapter 2, section 1 assessment

1. What are the characteristics of Matter?
2. Explain the difference between chemical and physical properties.
3. What is a pure substance?
4. Ice melts and liquid water freezes at 0°C , liquid water boils and water vapor condenses at 100°C . Is this an example of a chemical or physical property? Explain your answer.
5. When bread dough bakes, gasses are produced, creating the spaces (holes) in the bread. Is baking bread an example of a chemical or physical property? Explain your reasoning.
6. How are elements and compounds similar?
7. How are elements and compounds different?
8. Plants make a sugar compound with the chemical formula $\text{C}_6\text{H}_{12}\text{O}_6$. What elements make up this compound?
9. How does a heterogeneous mixture differ from a homogeneous mixture?
10. Explain why seawater is a mixture?
11. Suppose you stir a little baking soda into water until the water looks clear again. How could you prove to someone that the clear material is a solution, and not a compound?
12. Look at the following chemical formulas: H_2O and H_2O_2 . Do these formulas represent the same compound? Explain.

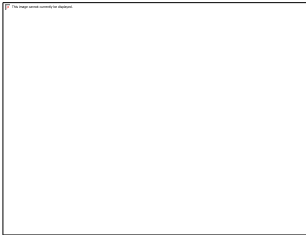
Name: _____ Due Date: _____ Period: _____

Chapter 2, section 2 assessment

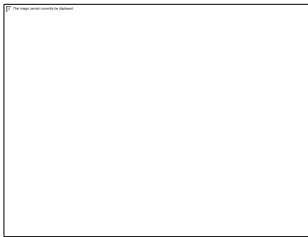
1. Define Mass
2. Why is mass more useful than weight for measuring matter?
3. Define volume
4. What measurements must you make to determine the density of a sample of matter?
5. How can you determine whether a solid substance is more or less dense than water?

Use the density triangle to complete the following density calculations.

6. A piece of metal has a volume of 38cm^3 and a mass of 277g . Calculate its density.



7. Iron has a density of 7.9 g/cm^3 . Calculate the mass of 38 cm^3 of Iron.



8. Lead has a density of 11.3 g/cm^3 . What would be the volume of 277 g . of Lead?

